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EN

Acts whose titles are printed in light type are those relating to day-to-day management of agricultural matters, and are generally valid for a limited period.

The titles of all other Acts are printed in bold type and preceded by an asterisk.

I

(Acts whose publication is obligatory)

COMMISSION DIRECTIVE 95/31/EC

of 5 July 1995

laying down specific criteria of purity concerning sweeteners for use in foodstuffs

(Text with EEA relevance)

THE COMMISSION OF THE EUROPEAN COMMUNITIES,

Having regard to the Treaty establishing the European Community,

Having regard to Council Directive 89/107/EEC of 21 December 1988 on the approximation of the laws of the Member States concerning food additives authorized for use in foodstuffs intended for human consumption ⁽¹⁾, as amended by Directive 94/34/EC ⁽²⁾, and in particular Article 3 (3) (a) thereof,

After consultation of the Scientific Committee on Food,

Whereas it is necessary to establish purity criteria for all sweeteners mentioned in European Parliament and Council Directive 94/35/EC of 30 June 1994 on sweeteners for use in foodstuffs ⁽³⁾;

Whereas it is necessary to take into account the specifications and analytical techniques for sweeteners as set out in the *Codex Alimentarius* and the Joint FAO/WHO Expert Committee on Food Additives (Jecfa);

Whereas food additives, prepared by production methods or starting materials significantly different from those included in the evaluation of the Scientific Committee for Food, or different from those mentioned in this Directive, should be submitted for evaluation by the Scientific Committee for Food with a view to full evaluation with emphasis on the purity criteria;

Whereas the measures provided for in this Directive are in line with the opinion of the Standing Committee on Foodstuffs,

HAS ADOPTED THIS DIRECTIVE:

Article 1

1. Purity criteria mentioned under Article 3 (3) (a) of Directive 89/107/EEC for sweeteners mentioned in Directive 94/35/EC are set out in the Annex.

2. The purity criteria for E 420 (i), E 420 (ii) and E 421 mentioned in the Annex to this Directive supersede the purity criteria for the said substances mentioned in the Annex to Council Directive 78/663/EEC ⁽⁴⁾.

Article 2

1. Member States shall bring into force the laws, regulations and administrative provisions necessary to comply with this Directive not later than 1 July 1996. They shall forthwith inform the Commission thereof.

When Member States adopt these provisions, these shall contain a reference to this Directive or shall be accompanied by such reference at the time of their official publication. The procedure for such reference shall be adopted by Member States.

⁽¹⁾ OJ No L 40, 11. 2. 1989, p. 27.

⁽²⁾ OJ No L 237, 10. 9. 1994, p. 1.

⁽³⁾ OJ No L 237, 10. 9. 1994, p. 3.

⁽⁴⁾ OJ No L 223, 14. 8. 1978, p. 7.

2. Products put on the market or labelled before that date which do not comply with this Directive may, however, be marketed until stocks are exhausted.

Article 3

This Directive shall enter into force on the 20th day following its publication in the *Official Journal of the European Communities*.

Article 4

This Directive is addressed to the Member States.

Done at Brussels, 5 July 1995.

For the Commission
Martin BANGEMANN
Member of the Commission

ANNEX

E 420 (i) — SORBITOL

Synonyms	D-glucitol, D-sorbitol
Definition	
<i>Chemical name</i>	D-glucitol
<i>Einecs</i>	200-061-5
<i>E number</i>	E 420 (i)
<i>Chemical formula</i>	$C_6H_{14}O_6$
<i>Relative molecular mass</i>	182,17
<i>Assay</i>	Content not less than 97% of total glycitols and not less than 91% of D-sorbitol on the dry weight basis. Glycitols are compounds with the structural formula $CH_2OH-(CHOH)_n-CH_2OH$, where 'n' is an integer
Description	White hygroscopic powder, crystalline powder, flakes or granules having a sweet taste
Identification	
A. <i>Solubility</i>	Very soluble in water, slightly soluble in ethanol
B. <i>Melting range</i>	88 to 102 °C
C. <i>Sorbitol monobenzylidene derivative</i>	To 5 g of the sample add 7 ml of methanol, 1 ml of benzaldehyde and 1 ml of hydrochloric acid. Mix and shake in a mechanical shaker until crystals appear. Filter with the aid of suction, dissolve the crystals in 20 ml of boiling water containing 1 g of sodium bicarbonate, filter while hot, cool the filtrate, filter with suction, wash with 5 ml of methanol-water mixture (1 in 2) and dry in air. The crystals so obtained melt between 173 and 179 °C
Purity	
<i>Water content</i>	Not more than 1% (Karl Fischer method)
<i>Sulphated ash</i>	Not more than 0,1% expressed on dry weight basis
<i>Reducing sugars</i>	Not more than 0,3% expressed as glucose on dry weight basis
<i>Total sugars</i>	Not more than 1% expressed as glucose on dry weight basis
<i>Chlorides</i>	Not more than 50 mg/kg expressed on dry weight basis
<i>Sulphates</i>	Not more than 100 mg/kg expressed on dry weight basis
<i>Nickel</i>	Not more than 2 mg/kg expressed on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis

<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis
E 420 (ii) — SORBITOL SYRUP	
Synonyms	D-clucitol syrup
Definition	
<i>Chemical name</i>	Sorbitol syrup formed by hydrogenation of glucose syrup is composed of D-sorbitol, D-mannitol and hydrogenated saccharides. The part of the product which is not D-sorbitol is composed mainly of hydrogenated oligosaccharides formed by the hydrogenation of glucose syrup used as raw material (in which case the syrup is non-crystallizing) or mannitol. Minor quantities of glycitols where $n \leq 4$ may be present. Glycitols are compounds with the structural formula $\text{CH}_2\text{OH}-(\text{CHOH})_n-\text{CH}_2\text{OH}$, where 'n' is an integer
<i>Einecs</i>	270-337-8
<i>E number</i>	E 420 (ii)
<i>Assay</i>	Content not less than 69% total solids and not less than 50% of D-sorbitol on the anhydrous basis
Description	Clear colourless and sweet tasting aqueous solution
Identification	
<i>A. Solubility</i>	Miscible with water, with glycerol, and with propane-1,2-diol
<i>B. Sorbitol monobenzylidene derivative</i>	To 5 g of the sample add 7 ml of methanol, 1 ml of benzaldehyde and 1 ml of hydrochloric acid. Mix and shake in a mechanical shaker until crystals appear. Filter with the aid of suction, dissolve the crystals in 20 ml of boiling water containing 1 g of sodium bicarbonate, filter while hot. Cool the filtrate filter with suction, wash with 5 ml of methanol-water mixture (1 in 2) and dry in air. The crystals so obtained melt between 173 and 179 °C
Purity	
<i>Water content</i>	Not more than 31% (Karl Fischer method)
<i>Sulphated ash</i>	Not more than 0,1% expressed on dry weight basis
<i>Reducing sugars</i>	Not more than 0,3% expressed as glucose on dry weight basis
<i>Chlorides</i>	Not more than 50 mg/kg expressed on dry weight basis
<i>Sulphates</i>	Not more than 100 mg/kg expressed on dry weight basis
<i>Nickel</i>	Not more than 2 mg/kg expressed on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis

E 421 — MANNITOL

Synonyms	D-mannitol
Definition	
<i>Chemical name</i>	D-mannitol
<i>Einecs</i>	200-711-8
<i>E number</i>	E 421
<i>Chemical formula</i>	C ₆ H ₁₄ O ₆
<i>Relative molecular mass</i>	182,2
<i>Assay</i>	Content not less than 96 % D-mannitol on the dried basis
Description	Sweet tasting, white, odourless, crystalline powder
Identification	
<i>A. Solubility</i>	Soluble
Purity	
<i>Loss on drying</i>	Not more than 0,3 % (105 °C, four hours)
<i>pH</i>	Between 5 and 8 Add 0,5 ml of a saturated solution of potassium chloride to 10 ml of a 10 % w/v solution of the sample then measure the pH
<i>Specific rotation</i>	$(\alpha)_{\text{D}}^{20}$ Specific rotation in a borated solution calculated with reference to the anhydrous substance is between + 23 and + 25°
<i>Sulphated ash</i>	Not more than 0,1 % expressed on dry weight basis
<i>Reducing sugars</i>	Not more than 0,3 % expressed as glucose on dry weight basis
<i>Total sugars</i>	Not more than 1 % expressed as glucose on dry weight basis
<i>Chlorides</i>	Not more than 70 mg/kg expressed on dry weight basis
<i>Sulphates</i>	Not more than 100 mg/kg expressed on dry weight basis
<i>Nickel</i>	Not more than 2 mg/kg expressed on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis

E 953 — ISOMALT

Synonyms	Hydrogenated isomaltulose, hydrogenated palatinose
Definition	
<i>Chemical name</i>	Isomalt is a mixture of: D-glucopyranosyl-1,6-D-glucitol and D-glucopyranosyl-1,1-D-mannitol dihydrate
<i>Einecs</i>	
<i>E number</i>	E 953
<i>Chemical formula</i>	D-glucopyranosyl-1,6-D-glucitol: $C_{12}H_{24}O_{11}$ D-glucopyranosyl-1,1-D-mannitol dihydrate: $C_{12}H_{24}O_{11} \cdot 2H_2O$
<i>Relative molecular mass</i>	D-glucopyranosyl-1,6-D-glucitol: 344,32 D-glucopyranosyl-1,1-D-mannitol dihydrate: 380,32
<i>Assay</i>	Content not less than 95 % of the mixture of D-glucopyranosyl-1,6-D-glucitol and D-glucopyranosyl-1,1-D-mannitol dihydrate determined on the anhydrous basis
Description	Odourless, white, sweet tasting, crystalline slightly hygroscopic substance
Identification	
A. <i>Solubility</i>	Slightly soluble in water, insoluble in ethanol
B. <i>Specific rotation</i>	$(\alpha)_D^{20}$: between + 90 and + 92° (4 % w/v solution)
C. <i>Melting range</i>	145 to 150 °C
Tests	
<i>Water content</i>	Not more than 7 % (Karl Fischer method)
<i>Sulphated ash</i>	Not more than 0,05 expressed on dry weight basis
<i>Reducing sugars</i>	Not more than 1,5 % expressed as glucose on dry weight basis
<i>Nickel</i>	Not more than 2 mg/kg expressed on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed on dry weight basis

E 965 (i) — MALTITOL

Synonyms	D-maltitol, hydrogenated maltose
Definition	
<i>Chemical name</i>	(α)-D-glucopyranosyl-1,4-D-glucitol

<i>Einecs</i>	209-567-0
<i>E number</i>	E 965 (i)
<i>Chemical formula</i>	$C_{12}H_{24}O_{11}$
<i>Relative molecular mass</i>	344,31
<i>Assay</i>	Content not less than 98 % D-mannitol $C_{12}H_{24}O_{11}$ on the anhydrous basis
Description	Sweet tasting, white crystalline powder
Identification	
A. <i>Solubility</i>	Very soluble in water, slightly soluble in ethanol
B. <i>Melting range</i>	148 to 151 °C
C. <i>Specific rotation</i>	$(\alpha)_D^{20} = + 105,5$ to $+ 105,5^\circ$ (5 % w/v solution)
Purity	
<i>Water content</i>	Not more than 1 % (Karl Fischer method)
<i>Sulphated ash</i>	Not more than 0,1 % expressed on dry weight basis
<i>Reducing sugars</i>	Not more than 0,1 % expressed as glucose on dry weight basis
<i>Chlorides</i>	Not more than 50 mg/kg expressed on dry weight basis
<i>Sulphates</i>	Not more than 100 mg/kg expressed on dry weight basis
<i>Nickel</i>	Not more than 2 mg/kg expressed on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis

E 965 (ii) — MALTITOL SYRUP

Synonyms	Hydrogenated high maltose-glucose syrup, hydrogenated glucose syrup
Definition	
<i>Chemical name</i>	A mixture consisting of mainly maltitol with sorbitol and hydrogenated oligo- and polysaccharides. It is manufactured by the catalytic hydrogenation of high maltose-content glucose syrup. The article of commerce is supplied both as a syrup and as a solid product
<i>Einecs</i>	270-337-8

<i>E number</i>	E 965 (ii)
<i>Assay</i>	The following ranges apply on the anhydrous basis: Maltitol not less than 50 % Sorbitol not more than 8 % Maltotriitol not more than 25 % Hydrogenated polysaccharides containing more than three glucose or glucitol units not more than 30 %
Description	Sweet-tasting, colourless and odourless, clear viscous liquids or sweet-tasting white crystalline masses
Identification	
<i>A. Solubility</i>	Very soluble in water, slightly soluble in ethanol
<i>B. Thin layer chromatography</i>	Examine by the thin layer chromatography using a plate coated with a 0,25 mm layer of chromatographic silica gel
Purity	
<i>Water content</i>	Not more than 31 % (Karl Fischer method)
<i>Sulphated ash</i>	Not more than 0,1 % expressed on dry weight basis
<i>Reducing sugars</i>	Not more than 0,3 % expressed as glucose on dry weight basis
<i>Chlorides</i>	Not more than 50 mg/kg expressed on dry weight basis
<i>Sulphates</i>	Not more than 100 mg/kg expressed on dry weight basis
<i>Nickel</i>	Not more than 2 mg/kg expressed on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis

E 966 — LACTITOL

Synonyms	Lactit, lactositol, lactobiosit
Definition	
<i>Chemical name</i>	4-O-β-D-galactopyranosyl-D-glucitol
<i>Einecs</i>	209-566-5
<i>E number</i>	E 966
<i>Chemical formula</i>	C ₁₂ H ₂₄ O ₁₁
<i>Relative molecular mass</i>	344,32
<i>Assay</i>	Not less than 95 % on the dry weight basis

Description	Sweet-tasting crystalline powders or colourless solutions. Crystalline products occur in anhydrous, monohydrate and dihydrate forms
Identification	
A. <i>Solubility</i>	Very soluble in water
B. <i>Specific rotation</i>	$(\alpha)_D^{20} = + 13$ to $+ 16^\circ$ calculated on the anhydrous basis (10 % w/v aqueous solution)
Purity	
<i>Water content</i>	Crystalline products; not more than 10,5 % (Karl Fischer method)
<i>Other polyols</i>	Not more than 2,5 % on the anhydrous basis
<i>Reducing sugars</i>	Not more than 0,2 % expressed as glucose on dry weight basis
<i>Chlorides</i>	Not more than 100 mg/kg expressed on dry weight basis
<i>Sulphates</i>	Not more than 200 mg/kg expressed on dry weight basis
<i>Sulphated ash</i>	Not more than 0,1 % expressed on dry weight basis
<i>Nickel</i>	Not more than 2 mg/kg expressed on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis

E 967 — XYLITOL

Synonyms	Xylitol
Definition	
<i>Chemical name</i>	D-xylitol
<i>Einecs</i>	201-788-0
<i>E number</i>	E 967
<i>Chemical formula</i>	$C_5H_{12}O_5$
<i>Relative molecular mass</i>	152,15
<i>Assay</i>	Not less than 98,5 % as xylitol on the anhydrous basis
Description	White, crystalline powder, practically odourless with a very sweet taste
Identification	
A. <i>Solubility</i>	Very soluble in water, sparingly soluble in ethanol
B. <i>Melting range</i>	92 to 96 °C
C. <i>pH</i>	5 to 7 (10 % w/v aqueous solution)

Purity	
<i>Loss on drying</i>	Not more than 0,5 %. Dry 0,5 g of sample in a vacuum over phosphorus at 60 °C for four hours
<i>Sulphated ash</i>	Not more than 0,1 % expressed on dry weight basis
<i>Reducing sugars</i>	Not more than 0,2 % expressed as glucose on dry weight basis
<i>Other polyhydric alcohols</i>	Not more than 1 % expressed on dry weight basis
<i>Nickel</i>	Not more than 2 mg/kg expressed on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis
<i>Chlorides</i>	Not more than 100 mg/kg expressed on dry weight basis
<i>Sulphates</i>	Not more than 200 mg/kg expressed on dry weight basis

E 950 — ACESULFAME K

Synonyms	Acesulfame potassium, acesulfam, potassium salt of 3,4-dihydro-6-methyl-1,2,3-oxathiazin-4-one-2,2-dioxide
Definition	
<i>Chemical name</i>	6-methyl-1,2,3-oxathiazin-4(3H)-one-2,2-dioxide potassium salt
<i>Einecs</i>	259-715-3
<i>E number</i>	E 950
<i>Chemical formula</i>	C ₄ H ₄ NO ₄ SK
<i>Relative molecular mass</i>	201,24
<i>Assay</i>	Not less than 99 % of C ₄ H ₄ NO ₄ SK on the anhydrous basis
Description	Odourless, white, crystalline powder having an intensively sweet taste. Approximately 200 times as sweet as sucrose
Identification	
A. <i>Solubility</i>	Very soluble in water, very slightly soluble in ethanol
B. <i>Ultra-violet absorption</i>	Maximum 227 ± 2 nm for a solution of 10 mg in 1 000 ml of water
Purity	
<i>Loss on drying</i>	Not more than 1 % (105 °C, two hours)

<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Selenium</i>	Not more than 30 mg/kg expressed on dry weight basis
<i>Fluoride</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis
E 951 — ASPARTAME	
Synonyms	Aspartyl phenylalanine methyl ester
Definition	
<i>Chemical name</i>	N-L- α -(Aspartyl-L-phenylalanine-1-methyl ester, 3-amino-N-(α -carbomethoxy-phenethyl)-succinamic acid-N-methyl ester
<i>Einecs</i>	245-261-3
<i>E number</i>	E 951
<i>Chemical formula</i>	$C_{14}H_{18}N_2O_5$
<i>Relative molecular mass</i>	294,31
<i>Assay</i>	Not less than 98 % and not more than 102 % of $C_{14}H_{18}N_2O_5$ on the anhydrous basis
Description	White, odourless, crystalline powder having a sweet taste. Approximately 200 times as sweet as sucrose
Identification	
<i>Solubility</i>	Slightly soluble in water and in ethanol
Purity	
<i>Loss on drying</i>	Not more than 4,5 % (105 °C, four hours)
<i>Sulphated ash</i>	Not more than 0,2 % expressed on dry weight basis
<i>pH</i>	Between 4,5 and 6,0 (1 in 125 solution)
<i>Transmittance</i>	The transmittance of a 1 % solution in 2N hydrochloric acid, determined in a 1-cm cell at 430 nm with a suitable spectrophotometer, using 2N hydrochloric acid as a reference, is not less than 0,95, equivalent to an absorbance of not more than approximately 0,022
<i>Specific rotation</i>	$(\alpha)_D^{20}$: +14,5 to +16,5° Determine in a 4 in 100/15 N formic acid solution within 30 minutes after preparation of the sample solution
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis

<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis
<i>5-Benzyl-3,6-dioxo-2-piperazineacetic acid</i>	Not more than 1,5% expressed on dry weight basis

E 952 — CYCLAMIC ACID AND ITS Na AND Ca SALTS

(I) CYCLAMIC ACID

Synonyms	Cyclohexylsulphamic acid, cyclamate
Definition	
<i>Chemical name</i>	Cyclohexanesulphamic acid, cyclohexylaminosulphonic acid
<i>Einecs</i>	202-898-1
<i>E number</i>	E 952
<i>Chemical formula</i>	$C_6H_{13}NO_3S$
<i>Relative molecular mass</i>	179,24
<i>Assay</i>	Cyclohexylsulphamic acid contains not less than 98% and not more than the equivalent of 102% of $C_6H_{13}NO_3S$, calculated on the anhydrous basis
Description	A practically colourless, white crystalline powder with a sweet-sour taste. Approximately 40 times as sweet as sucrose
Identification	
A. <i>Solubility</i>	Soluble in water and in ethanol
B. <i>Precipitation test</i>	Acidify a 2% solution with hydrochloric acid, add 1 ml of an approximately molar solution of barium chloride in water and filter if any haze or precipitate forms. To the clear solution add 1 ml of a 10% solution of sodium nitrite. A white precipitate forms.
Purity	
<i>Loss on drying</i>	Not more than 1% (105 °C, one hour)
<i>Selenium</i>	Not more than 30 mg/kg expressed as selenium on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Cyclohexylamine</i>	Not more than 10 mg/kg expressed on dry weight basis
<i>Dicyclohexylamine</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Aniline</i>	Not more than 1 mg/kg expressed on dry weight basis

(II) SODIUM CYCLAMATE

Synonyms	Cyclamate, sodium salt of cyclamic acid
Definition	
<i>Chemical name</i>	Sodium cyclohexanesulphamate, sodium cyclohexylsulphamate
<i>Einecs</i>	205-348-9
<i>E number</i>	E 952
<i>Chemical formula</i>	$C_6H_{12}NNaO_3S$ and the dihydrate form $C_6H_{12}NNaO_3S \cdot 2H_2O$
<i>Relative molecular mass</i>	201,22 calculated on the anhydrous form 237,22 calculated on the hydrated form
<i>Assay</i>	Not less than 98 % and not more than 102 % on the dried basis Dihydrate form: not less than 84 % on the dried basis
Description	White, odourless crystals or crystalline powder. Approximately 30 times as sweet as sucrose
Identification	
<i>Solubility</i>	Soluble in water, practically insoluble in ethanol
Purity	
<i>Loss on drying</i>	Not more than 1 % (105 °C, one hour) Not more than 15,2 % (105 °C, two hours) for the dihydrate form
<i>Selenium</i>	Not more than 30 mg/kg expressed as selenium on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis
<i>Cyclohexylamine</i>	Not more than 10 mg/kg expressed on dry weight basis
<i>Dicyclohexylamine</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Aniline</i>	Not more than 1 mg/kg expressed on dry weight basis

(III) CALCIUM CYCLAMATE

Synonyms	Cyclamate, calcium salt of cyclamic acid
Definition	
<i>Chemical name</i>	Calcium cyclohexanesulphamate, calcium cyclohexylsulphamate
<i>Einecs</i>	205-349-4
<i>E number</i>	E 952
<i>Chemical formula</i>	$C_{12}H_{24}CaN_2O_6S_2 \cdot 2H_2O$

<i>Relative molecular mass</i>	432,57
<i>Assay</i>	Not less than 98 % and not more than 10 % on the dried basis
Description	White, colourless crystals or crystalline powder. Approximately 30 times as sweet as sucrose
Identification	
<i>Solubility</i>	Soluble in water, sparingly soluble in ethanol
Purity	
<i>Loss on drying</i>	Not more than 1 % (105 °C, one hour) Not more than 8,5 % (140 °C, four hours) for the dihydrate form
<i>Selenium</i>	Not more than 30 mg/kg expressed as selenium on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis
<i>Cyclohexylamine</i>	Not more than 10 mg/kg expressed on dry weight basis
<i>Dicyclohexylamine</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Aniline</i>	Not more than 1 mg/kg expressed on dry weight basis

E 954 — SACCHARIN AND ITS Na. K AND Ca SALTS

(I) SACCHARIN

Definition

<i>Chemical name</i>	3-oxo-2,3dihydrobenzo(d)isothiazol-1,1-dioxide
<i>Einecs</i>	201-321-0
<i>E number</i>	E 954
<i>Chemical formula</i>	C ₇ H ₅ NO ₃ S
<i>Relative molecular mass</i>	183,18
<i>Assay</i>	Not less than 99 % and not more than 101,0 % of C ₇ H ₅ NO ₃ S on the anhydrous basis
Description	White crystals or a white crystalline powder, odourless or with a faint, aromatic odour having a sweet taste even in very dilute solutions. Approximately between 300 and 500 times as sweet as sucrose
Identification	
<i>Solubility</i>	Slightly soluble in water, soluble in basic solutions, sparingly soluble in ethanol

Purity	
<i>Loss on drying</i>	Not more than 1% (105 °C, two hours)
<i>Melting range</i>	226 to 230 °C
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Selenium</i>	Not more than 30 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis
<i>Sulphated ash</i>	Not more than 0,2% expressed on dry weight basis
<i>Benzoic and salicylic acid</i>	To 10 ml of a 1 in 20 solution, previously acidified with five drops of acetic acid, add three drops of an approximately molar solution of ferric chloride in water. No precipitate or violet colour appears
<i>o-Toluenesulphonamide</i>	Not more than 10 mg/kg expressed on dry weight basis
<i>p-Toluenesulphonamide</i>	Not more than 10 mg/kg expressed on dry weight basis
<i>Benzoic acid p-sulfonamide</i>	Not more than 25 mg/kg expressed on dry weight basis
<i>Readily carbonizable substances</i>	Absent
(II) SODIUM SACCHARIN	
Synonyms	Saccharin, sodium salt of saccharin
Definition	
<i>Chemical name</i>	Sodium o-benzosulphimide, sodium salt of 2,3-dihydro-3-oxobenzisulfonazole, oxobenzisulfonazole, 1,2-benzisothiazolin-3-one-1, 1-dioxide sodium salt dihydrate
<i>Einecs</i>	204-886-1
<i>E number</i>	E 954
<i>Chemical formula</i>	$C_7H_4NNaO_3S \cdot 2H_2O$
<i>Relative molecular mass</i>	241,19
<i>Assay</i>	Not less than 99% and not more than 101% of $C_7H_4NNaO_3S$ on the anhydrous basis
Description	White crystals or a white crystalline efflorescent powder, odourless or with a faint, odour, having an intensely sweet taste, even in very dilute solutions. Approximately between 300 and 500 times as sweet as sucrose in dilute solutions
Identification	
<i>Solubility</i>	Freely soluble in water, sparingly soluble in ethanol
Purity	
<i>Loss on drying</i>	Not more than 15% (120 °C, four hours)

<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Selenium</i>	Not more than 30 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis
<i>Benzoic and salicylic acid</i>	To 10 ml of a 1 in 20 solution, previously acidified with five drops of acetic acid, add three drops of an approximately molar solution of ferric chloride in water. No precipitate or violet colour appears
<i>o-Toluenesulphonamide</i>	Not more than 10 mg/kg expressed on dry weight basis
<i>p-Toluenesulphonamide</i>	Not more than 10 mg/kg expressed on dry weight basis
<i>Benzoic acid p-sulfonamide</i>	Not more than 25 mg/kg expressed on dry weight basis
<i>Readily carbonizable substances</i>	Absent
(III) CALCIUM SACCHARIN	
Synonyms	Saccharin, calcium salt of saccharin
Definition	
<i>Chemical name</i>	Calcium o-benzosulphimide, calcium salt of 2,3-dihydro-3-oxobenzisulphonazole, 1,2-benzisothiazolin-3-one-1,1-dioxide calcium salt hydrate (2:7)
<i>Einecs</i>	229-349-0
<i>E number</i>	E 954
<i>Chemical formula</i>	$C_{14}H_8CaN_2O_6S_2 \cdot 3\frac{1}{2}H_2O$
<i>Relative molecular mass</i>	467,48
<i>Assay</i>	Not less than 95 % of $C_{14}H_8CaN_2O_6S_2$ on the anhydrous basis
Description	White crystals or a white crystalline powder, odourless or with a faint odour, having an intensely sweet taste, even in very dilute solutions. Approximately between 300 and 500 times as sweet as sucrose in dilute solutions
Identification	
<i>Solubility</i>	Freely soluble in water, soluble in ethanol
Purity	
<i>Loss on drying</i>	Not more than 13,5 % (120 °C, four hours)
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Selenium</i>	Not more than 30 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis

<i>Benzoic and salicylic acid</i>	To 10 ml of a 1 in 20 solution, previously acidified with five drops of acetic acid, add three drops of an approximately molar solution of ferric chloride in water. No precipitate or violet colour appears
<i>o-Toluenesulphonamide</i>	Not more than 10 mg/kg expressed on dry weight basis
<i>p-Toluenesulphonamide</i>	Not more than 10 mg/kg expressed on dry weight basis
<i>Benzoic acid p-sulfonamide</i>	Not more than 25 mg/kg expressed on dry weight basis
<i>Readily carbonizable substances</i>	Absent
(IV) POTASSIUM SACCHARIN	
Synonyms	Saccharin, potassium salt of saccharin
Definition	
<i>Chemical name</i>	Potassium o-benzosulphimide, potassium salt of 2,3-dihydro-3-oxobenzisulphonazole, potassium salt of 1,2-benzisothiazolin-3-one-1,1-dioxide monohydrate
<i>Einecs</i>	
<i>E number</i>	E 954
<i>Chemical formula</i>	$C_7H_4KNO_3S \cdot H_2O$
<i>Relative molecular mass</i>	239,77
<i>Assay</i>	Not less than 99 % and not more than 101 % of $C_7H_4KNO_3S$ on the anhydrous basis
Description	White crystals or a white crystalline powder, odourless or with a faint odour, having an intensely sweet taste, even in very dilute solutions. Approximately between 300 and 500 times as sweet as sucrose
Identification	
<i>Solubility</i>	Freely soluble in water, sparingly soluble in ethanol
Purity	
<i>Loss on drying</i>	Not more than 8 % (120 °C, four hours)
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Selenium</i>	Not more than 30 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 1 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis
<i>Benzoic and salicylic acid</i>	To 10 ml of a 1 in 20 solution, previously acidified with five drops of acetic acid, add three drops of an approximately molar solution of ferric chloride in water. No precipitate or violet colour appears
<i>o-Toluenesulphonamide</i>	Not more than 10 mg/kg expressed on dry weight basis
<i>p-Toluenesulphonamide</i>	Not more than 10 mg/kg expressed on dry weight basis

<i>Benzoic acid p-sulfonamide</i>	Not more than 25 mg/kg expressed on dry weight basis
<i>Readily carbonizable substances</i>	Absent

E 957 — THAUMATIN

Synonyms	
Definition	
<i>Chemical name</i>	Thaumatocin is obtained by aqueous extraction (pH 2,5 to 4) of the arils of the fruit of the natural strain of <i>Thaumatococcus daniellii</i> (Benth) and consists essentially of the proteins thaumatocin I and thaumatocin II together with minor amounts of plant constituents derived from the source material
<i>Einecs</i>	258-822-2
<i>E number</i>	E 957
<i>Chemical formula</i>	Polypeptide of 207 aminoacids
<i>Relative molecular mass</i>	Thaumatocin I 22209 Thaumatocin II 22293
<i>Assay</i>	Not less than 16 % nitrogen on the dried basis equivalent to not less than 94 % proteins (N × 5,8)
Description	Odourless, cream-coloured powder with an intensely sweet taste. Approximately 2 000 to 3 000 times as sweet as sucrose
Identification	
<i>Solubility</i>	Very soluble in water, insoluble in acetone
Purity	
<i>Loss on drying</i>	Not more than 9 % (105 °C to constant weight)
<i>Carbohydrates</i>	Not more than 3 % expressed on dry weight basis
<i>Sulphated ash</i>	Not more than 2 % expressed on dry weight basis
<i>Aluminium</i>	Not more than 100 mg/kg expressed on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	3 mg/kg expressed on dry weight basis
<i>Microbiological criteria</i>	Total aerobic microbial count: Max 1 000/g <i>E. Coli</i> : absent in 1 g

E 959 — NEOHESPERIDINE DIHYDROCHALCONE

Synonyms	Neohesperidin dihydrochalcone, NHDC, hesperetin dihydrochalcone-4'- β -neohesperidoside, neohesperidin DC
Definition	
<i>Chemical name</i>	2-O- α -L-rhamnopyranosyl-4'- β -D-glucopyranosyl hesperetin dihydrochalcone; obtained by catalytic hydrogenation of neohesperidin

<i>Einecs</i>	243-978-6
<i>E number</i>	E 959
<i>Chemical formula</i>	$C_{28}H_{36}O_{15}$
<i>Relative molecular mass</i>	612,6
<i>Assay</i>	Content not less than 96 % on the dried basis
Description	Off white, odourless, crystalline powder having a characteristic, intensive sweet taste. Approximately between 1 000 and 1 800 times as sweet as sucrose
Identification	
<i>A. Solubility</i>	Freely soluble in hot water, very slightly soluble in cold water, practically insoluble in ether and benzene
<i>B. Ultraviolet absorption maximum</i>	282 to 283 nm for a solution of 2 mg in 100 ml methanol
<i>C. Neu's test</i>	Dissolve about 10 mg of neohesperidine DC in 1 ml methanol, add 1 ml of a 1 % 2-aminoethyl diphenyl borate methanolic solution. A bright yellow colour is produced
Purity	
<i>Loss on drying</i>	Not more than 11 % (105 °C, three hours)
<i>Sulphated ash</i>	Not more than 0,2 % expressed on dry weight basis
<i>Arsenic</i>	Not more than 3 mg/kg expressed on dry weight basis
<i>Lead</i>	Not more than 2 mg/kg expressed on dry weight basis
<i>Heavy metals</i>	Not more than 10 mg/kg expressed as Pb on dry weight basis

SIXTH COMMISSION DIRECTIVE 95/32/EC

of 7 July 1995

relating to methods of analysis necessary for checking the composition of cosmetic products

(Text with EEA relevance)

THE COMMISSION OF THE EUROPEAN COMMUNITIES,

Having regard to the Treaty establishing the European Community,

Having regard to Council Directive 76/768/EEC of 27 July 1976 on the approximation of the laws of the Member States relating to cosmetic products⁽¹⁾, as last amended by Commission Directive 94/32/EC⁽²⁾, and in particular Article 8 (1) thereof,

Whereas Directive 76/768/EEC provides for the official testing of cosmetic products with the aim of ensuring that the conditions laid down by Commission provisions concerning the composition of cosmetic products are satisfied;

Whereas all the necessary methods of analysis should be laid down as quickly as possible; whereas certain methods have already been adopted in Commission Directives 80/1335/EEC⁽³⁾, as amended by Directive 87/143/EEC⁽⁴⁾, 82/434/EEC⁽⁵⁾, as amended by Directive 90/207/EEC⁽⁶⁾, 83/514/EEC⁽⁷⁾, 85/490/EEC⁽⁸⁾ and 93/73/EEC⁽⁹⁾;

Whereas the identification and determination of benzoic acid, 4-hydroxybenzoic acid, sorbic acid, salicylic acid and propionic acid in cosmetic products and the identification and determination of hydroquinone, hydroquinone monomethylether, hydroquinone monoethylether and hydroquinone monobenzylether in cosmetic products constitute a sixth step;

Whereas the measures provided for in this Directive are in accordance with the opinion of the Committee on the adaptation of Directive 76/768/EEC to technical progress,

HAS ADOPTED THIS DIRECTIVE:

Article 1

Member States shall take all the necessary steps to ensure that during official testing of cosmetic products:

— identification and determination of benzoic acid, 4-hydroxybenzoic acid, sorbic acid, salicylic acid and propionic acid,

— identification and determination of hydroquinone, hydroquinone monomethylether, hydroquinone monoethylether and hydroquinone monobenzylether,

shall be carried out in accordance with the methods described in the Annex.

Article 2

1. Member States shall bring into force the laws, regulations and administrative provisions necessary to comply with this Directive no later than 30 September 1996. They shall forthwith inform the Commission thereof.

When Member States adopt these provisions, these shall contain a reference to this Directive or shall be accompanied by such reference at the time of their official publication. The procedure for such reference shall be adopted by Member States.

2. Member States shall communicate to the Commission the provisions of national law which they adopt in the field covered by this Directive.

*Article 3*This Directive shall enter into force on the 20th day following its publication in the *Official Journal of the European Communities*.*Article 4*

This Directive is addressed to the Member States.

Done at Brussels, 7 July 1995.

For the Commission

Emma BONINO

Member of the Commission

(¹) OJ No L 262, 27. 9. 1976, p. 169.

(²) OJ No L 181, 15. 7. 1994, p. 31.

(³) OJ No L 383, 31. 12. 1980, p. 27.

(⁴) OJ No L 57, 27. 2. 1987, p. 56.

(⁵) OJ No L 185, 30. 6. 1982, p. 1.

(⁶) OJ No L 108, 28. 4. 1990, p. 92.

(⁷) OJ No L 291, 24. 10. 1983, p. 9.

(⁸) OJ No L 295, 7. 11. 1985, p. 30.

(⁹) OJ No L 231, 14. 9. 1993, p. 34.

ANNEX

I. IDENTIFICATION AND DETERMINATION OF BENZOIC ACID, 4-HYDROXYBENZOIC ACID, SORBIC ACID, SALICYLIC ACID AND PROPIONIC ACID IN COSMETIC PRODUCTS**1. Scope and field of application**

The method is applicable to the identification and determination of benzoic acid, 4-hydroxybenzoic acid, sorbic acid, salicylic acid and propionic acid in cosmetic products. Separate procedures describe the identification of these preservatives; the determination of propionic acid; and the determination of 4-hydroxybenzoic acid, salicylic acid, sorbic acid and benzoic acid.

2. Definition

The amounts of benzoic acid, 4-hydroxybenzoic acid, salicylic acid, sorbic acid and propionic acid determined by this method are expressed as percentage by mass of the free acids.

A. IDENTIFICATION**1. Principle**

Following acid/base extraction of the preservatives, the extract is analysed by thin layer chromatography (TLC) employing on-date derivatization. Depending on the results, the identification is confirmed by high performance liquid chromatography (HPLC) or, in the case of propionic acid, by gas chromatography (GC).

2. Reagents**2.1. General**

All reagents must be of analytical purity. Water used must be distilled water, or water of at least equivalent purity

2.2. Acetone**2.3. Diethyl ether****2.4. Acetonitrile****2.5. Toluene****2.6. n-Hexane****2.7. Paraffin, liquid****2.8. Hydrochloric acid, 4 M****2.9. Potassium hydroxide, aqueous, 4 M****2.10. Calcium chloride, CaCl₂·2H₂O****2.11. Lithium carbonate, Li₂CO₃****2.12. 2-Bromo-2'-acetonaphthone****2.13. 4-Hydroxybenzoic acid****2.14. Salicylic acid****2.15. Benzoic acid****2.16. Sorbic acid****2.17. Propionic acid**

- 2.18. Reference solutions
Prepare 0,1 % (m/v) solutions (100 mg/100 ml) of each of the five preservatives (2.13 to 2.17) in diethyl ether
- 2.19. Derivatization reagent
0,5 % (m/v) solution of 2-bromo-2'-acetonaphthone (2.12) in acetonitrile (2.4) (50 mg/10 ml). This solution should be prepared weekly and stored in a refrigerator
- 2.20. Catalyst solution
0,3 % (m/v) solution of lithium carbonate (2.11) in water (300 mg/100 ml). This solution should be freshly prepared
- 2.21. Development solvent
Toluene (2.5)/Acetone (2.2) (20:0,5, v/v)
- 2.22. Liquid paraffin (2.7)/n-hexane (2.6) (1:2, v/v)

3. Apparatus

Ordinary laboratory equipment

- 3.1. Water bath, capable of maintaining a temperature of 60 °C
- 3.2. Developing tank
- 3.3. Ultraviolet light source, 254 and 366 nm
- 3.4. Thin layer plates, Kieselgel 60, without fluorescence indicator, 20 × 20 cm, layer thickness 0,25 mm with concentrating zone 2,5 × 20 cm (Merck 11845, or equivalent)
- 3.5. Microsyringe, 10 µl
- 3.6. Microsyringe, 25 µl
- 3.7. Oven, capable of maintaining temperatures up to 105 °C
- 3.8. 50-ml glass tubes with screw cap
- 3.9. Filter paper, diameter 90 mm, Schleicher & Schull, Weissband No 5892, or equivalent
- 3.10. Universal pH indicator paper, pH 1—11
- 3.11. 5-ml glass sample vials
- 3.12. Rotating film evaporator (Rotavapor or equivalent)
- 3.13. Hot plate

4. Procedure

4.1. Sample preparation

Weigh approximately 1 g of the sample into a 50-ml glass tube with screw cap (3.8). Add four drops of hydrochloric acid 4 M (2.8) and 40 ml acetone (2.2). For strongly basic products such as toilet soap, 20 drops of hydrochloric acid 4 M (2.8) should be added. Check that the pH is approximately two, using indicator paper (3.10). Close the tube and shake vigorously for one minute.

If necessary to facilitate the extraction of the preservatives into the acetone phase, heat the mixture gently to about 60 °C to melt the liquid phase.

Cool the solution to room temperature and filter through a filter paper (3.9) into a conical flask.

Transfer 20 ml of the filtrate to a 200-ml conical flask, add 20 ml water and mix. Adjust the pH of the mixture to approximately 10 with potassium hydroxide 4 M (2.9), using indicator paper (3.10) to measure the pH.

Add 1 g calcium chloride (2.10) and shake vigorously. Filter through a filter paper (3.9) into a 250-ml separating funnel containing 75 ml diethyl ether (2.3) and shake vigorously for one minute. Allow to separate and draw off the aqueous layer into a 250 ml conical flask. Discard the ether layer. Using indicator paper (3.10), adjust the pH of the aqueous solution to approximately two with hydrochloric acid 4 M (2.8). Add 10 ml diethyl ether (2.3), stopper the flask and shake vigorously for one minute. allow to separate and transfer the ether layer to a rotating film evaporator (3.12). Discard the aqueous layer.

Evaporate the ether layer almost to dryness and redissolve the residue in 1 ml of diethyl ether (2.3). Transfer the solution to a sample vial (3.11).

4.2. Thin layer chromatography

For each of the references and the samples to be chromatographed, apply approximately 3 μ l lithium carbonate solution (2.20) and with a syringe (3.5) at equal distances on the start line in the concentration zone of a TLC plate (3.4) and dry in a stream of cold air.

Transfer the TLC plate to a hot plate (3.13), heated at 40 °C, in order to keep the spots as small as possible. With a microsyringe (3.5) apply 10 μ l of each of the reference solutions (2.18) and the sample solution (4.1) to the start line of the plate, on the exact spots where the lithium carbonate solution has been applied.

Finally apply approximately 15 μ l derivatization reagent (2.19) (2-bromo-2'-acetonaphthone solution), again on the exact spots where the reference/sample solutions and the lithium carbonate solution have been applied.

Heat the TLC plate in an oven (3.7) at 80 °C for 45 minutes. After cooling, develop the plate in a tank (3.2), that has equilibrated for 15 minutes (without the use of filter paper lining), using development solvent 2.21 (toluene/acetone), until the solvent front has reached a distance of 15 cm (this may take approximately 80 minutes).

Dry the plate in a stream of cold air and examine the spots obtained under UV light (3.3). To enhance the fluorescence of the weak spots, the TLC plate may be dipped in liquid paraffin/n-hexane (2.22).

5. Identification

Calculate the R_f for each spot.

Compare the R_f and the behaviour under UV radiation obtained for the sample with that obtained for the reference solutions.

Draw a preliminary conclusion about the presence and identity of the preservatives present. Perform the HPLC described in Section B, or, when it appears that propionic acid is present, the GC described in Section C. Compare the retention times obtained with those of the reference solutions.

Combine the results from the TLC and HPLC or GC and base the final identification of the preservatives present in the sample on the combined results.

B. DETERMINATION OF BENZOIC ACID, 4-HYDROXYBENZOIC ACID, SORBIC ACID AND SALICYLIC ACID

1. Principle

After acidification, the sample is extracted with a mixture of ethanol and water. Following filtration the preservatives are determined by high performance liquid chromatography (HPLC).

2. Reagents

2.1. All reagents must be of analytical purity, and suitable for HPLC where appropriate. Water used must be distilled water, or water of at least equivalent purity

2.2. Ethanol, absolute

2.3. 4-Hydroxybenzoic acid

- 2.4. Salicylic acid
 - 2.5. Benzoic acid
 - 2.6. Sorbic acid
 - 2.7. Sodium acetate, (CH₃COONa.3H₂O)
 - 2.8. Acetic acid, (α)₄²⁰= 1,05 g/ml
 - 2.9. Acetonitrile
 - 2.10. Sulfuric acid, 2 M
 - 2.11. Potassium hydroxide, aqueous, 0,2 M
 - 2.12. 2-Methoxybenzoic acid
 - 2.13. Ethanol/water mixture
Mix nine volumes of ethanol (2.2) and one volume of water (2:1)
 - 2.14. Internal standard solution
Prepare a solution containing approximately 1 g 2-methoxybenzoic acid (2.12) in 500 ml ethanol/water mixture (2.13)
 - 2.15. Mobile phase for HPLC
 - 2.15.1. Acetate buffer: to 1 l of water add 6,35 g sodium acetate (2.7) and 20,0 ml acetic acid (2.8) and mix
 - 2.15.2. Prepare the mobile phase by mixing nine volumes acetate buffer (2.15.1) and one volume acetonitrile (2.9)
 - 2.16. Preservative stock solution
Accurately weigh approximately 0,05 g 4-hydroxybenzoic acid (2.3), 0,2 g salicylic acid (2.4), 0,2 g benzoic acid (2.5) and 0,05 g sorbic acid (2.6) in a 50-ml volumetric flask and make up to volume with ethanol/water mixture (2.13). Store this solution in a refrigerator. The solution is stable for one week
 - 2.17. Standard preservative solutions
Transfer respectively 8,00, 4,00, 2,00, 1,00 and 0,50 ml of the stock solution (2.16) into a series of 20-ml volumetric flasks. To each flask, add 10,00 ml internal standard solution (2.14) and 0,5 ml sulfuric acid 2 M (2.10). Make up to volume with ethanol/water mixture (2.13). These solutions must be freshly prepared
3. **Apparatus**
- Usual laboratory equipment not otherwise specified, and:
- 3.1. Water bath, set at 60 °C
 - 3.2. High performance liquid chromatograph with variable-wavelength UV detector and 10- μ l injection loop
 - 3.3. Analytical column
Stainless steel, length 12,5 to 25 cm, internal diameter 4,6 mm, packed with Nucleosil 5C18, or equivalent
 - 3.4. Filter paper, diameter: 90 mm, Schleicher and Schull, Weissband No 5892, or equivalent
 - 3.5. 50-ml glass tubes with screw cap

- 3.6. 5-ml glass sample vials
- 3.7. Boiling chips, carborundum, size 2 to 4 mm, or equivalent

4. Procedure

4.1. Sample preparation

4.1.1. Sample preparation without addition of internal standard

Weigh 1 g of the sample in a 50-ml glass tube with screw cap (3.5). Pipette 1,00 ml sulfuric acid 2 M (2.10) and 40,0 ml ethanol/water mixture (2.13) into the tube. Add approximately 1 g of boiling chips (3.7), close the tube and shake vigorously for at least one minute until a homogeneous suspension is obtained. To facilitate the extraction of the preservatives into the ethanol phase, place the tube for exactly five minutes in a water bath (3.1) kept at 60 °C.

Cool the tube immediately in a stream of cold water and store the extract at 5 °C for one hour.

Filter the extract through a filter paper (3.4). Transfer approximately 2 ml of the extract into a sample vial (3.6). Store the extract at 5 °C and perform the HPLC determination within 24 hours of preparation.

4.1.2. Sample preparation including addition of internal standard

Weigh to three decimal places $1 \pm 0,1$ g (a grams) of the sample in a 50-ml glass tube with screw cap (3.5). Add by pipette 1,00 ml sulfuric acid 2 M (2.10) and 30,0 ml ethanol/water mixture (2.13). Add approximately 1 g of boiling chips (3.7) and 10,00 ml internal standard solution (2.14). Close the tube and shake vigorously for at least one minute until an homogeneous suspension is obtained. To facilitate extraction of the preservatives into the ethanol phase place the tube for exactly five minutes in a water bath (3.1) kept at 60 °C.

Cool the tube immediately in a stream of cold water and store the extract at 5 °C for one hour.

Filter the extract through a filter paper (3.4). Transfer approximately 2 ml of the filtrate into a sample vial (3.6). Store the filtrate at 5 °C and perform the HPLC-determination within 24 hours of preparation.

4.2. High performance liquid chromatography

Mobile phase: acetonitrile/acetate buffer (2.15).

Adjust the flow rate of mobile phase through the column to $2,0 \pm 0,5$ ml/minute. Set the detector wavelength to 240 nm.

4.2.1. Calibration

Inject 10 μ l portions of each of the standard preservative solutions (2.17) into the liquid chromatography (3.2). For each solution determine the ratios of the peak heights of the investigated preservatives to the height of the internal standard peak obtained from the chromatograms. Plot a graph for each preservative relating the peak height ratio to the concentration of each standard solution.

Ascertain that a linear response is obtained for the standard solutions in the calibration procedure.

4.2.2. Determination

Inject 10 μ l of the sample extract (4.1.1) into the liquid chromatograph (3.2) and record the chromatogram. Inject 10 μ l of a standard preservative solution (2.17) and record the chromatogram. Compare the chromatograms obtained. If in the chromatogram of the sample extract (4.1.1) no peak appears to be present having approximately the same retention time as 2-methoxybenzoic acid (recommended internal standard), inject 10 μ l sample extract with added internal standard (4.1.2) into the liquid chromatograph and record the chromatogram.

If an interfering peak is observed in the chromatogram of the sample extract (4.1.1) having the same retention time as 2-methoxybenzoic acid, another appropriate internal standard should be selected. (If one of the preservatives under investigation is absent from the chromatogram, this preservative can be used as the internal standard.)

Ascertain whether the chromatograms obtained for a standard solution and the sample solution meet the following requirements:

- the peak separation of the worst separated pair shall be at least 0,90. (For definition of peak separation, see figure 1).

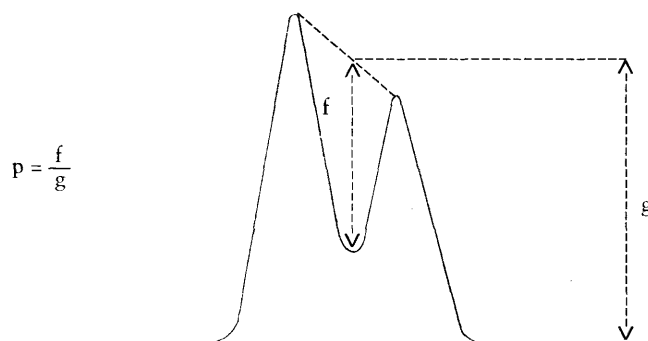


Figure 1: Peak separation

If the required separation is not achieved, either a more efficient column should be used, or the mobile phase composition should be adjusted until the requirement is met.

- The asymmetry factor A_s of all peaks obtained shall range between 0,9 to 1,5. (For definition of the peak asymmetry factor, see Figure 2). To record the chromatogram for the determination of the asymmetry factor a chart speed of at least 2 cm/minute is recommended.

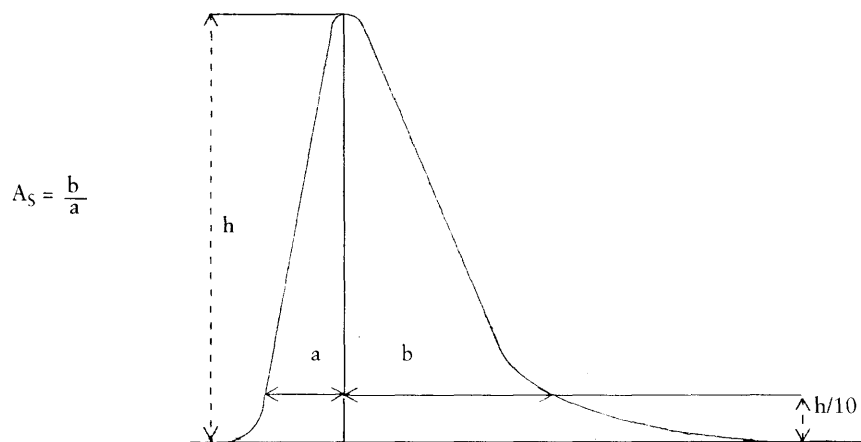


Figure 2: Peak asymmetry factor

- A steady baseline shall be obtained.

5. Calculation

Use the ratios of the heights of the peaks of the investigated preservatives to the height of the 2-methoxybenzoic acid (internal standard) peak and the calibration graph to calculate the concentration of the acid preservatives in the sample solution. Calculate the benzoic acid, 4-hydroxybenzoic acid, sorbic acid or salicylic acid content of the sample, as a percentage by mass (x_i), using the formula:

$$x_i \% (m/m) = \frac{100 \cdot 20 \cdot b}{10^6 \cdot a} = \frac{b}{500 \cdot a}$$

in which:

a = the mass (g) of the test portion (4.1.2),

b = the concentration ($\mu\text{g/ml}$) of the preservative in the sample extract (4.1.2) obtained from the calibration graph.

6. Repeatability ⁽¹⁾

For a 4-hydroxybenzoic acid content of 0,40 % the difference between the results of two determinations in parallel carried out on the same sample should not exceed an absolute value of 0,035 %.

For a benzoic acid content of 0,50 % the difference between the results of two determinations in parallel carried out on the same sample should not exceed an absolute value of 0,050 %.

For a salicylic acid content of 0,50 % the difference between the results of two determinations in parallel carried out on the same sample should not exceed an absolute value of 0,045 %.

For a sorbic acid content of 0,60 % the difference between the results of two determinations in parallel carried out on the same sample should not exceed an absolute value of 0,035 %.

7. Remarks

7.1. Results of a ruggedness test performed on the method indicated that the amount of sulfuric acid added to extract the acids from the sample is critical, and the limits for the amount of sample worked up should be kept within the prescribed boundaries.

7.2. If desired, an appropriate guard column can be used.

C. DETERMINATION OF PROPIONIC ACID**1. Scope and field of application**

This method is suitable for the determination of propionic acid, maximum concentration 2 % (m/m) in cosmetic products.

2. Definition

The concentration of propionic acid measured by this method is expressed as a percentage by mass (% m/m) of the product.

3. Principle

Following extraction of propionic acid from the product determination is performed by means of gas chromatography with the use of 2-methylpropionic acid as internal standard.

4. Reagents

All the reagents must be of analytical purity; distilled water or water of equivalent quality must be used.

4.1. Ethanol 96 % (v/v)

4.2. Propionic acid

4.3. 2-Methylpropionic acid

4.4. Orthophosphoric acid, 10 % (m/v)

4.5. Propionic acid solution

Accurately weigh approximately 1,00 g (p grams) of propionic acid into a 50-ml volumetric flask and make up to volume with ethanol (4.1)

4.6. Internal standard solution

Accurately weigh approximately 1,00 g (e grams) of 2-methylpropionic acid into a 50-ml volumetric flask and make up to volume with ethanol (4.1)

⁽¹⁾ ISO 5725.

5. Apparatus

- 5.1. Usual laboratory equipment, and:
- 5.2. Gas chromatograph with flame-ionization detector
- 5.3. Glass tube (20 × 150 mm) with screw cap
- 5.4. Water bath at 60 °C
- 5.5. 10 ml glass syringe with filter membrane (pore diameter: 0,45 µm)

6. Procedure**6.1. Sample preparation****6.1.1. Sample preparation without the internal standard**

Into a glass tube (5.3), weigh approximately 1 g of the sample. Add 0,5 ml of phosphoric acid (4.4) and 9,5 ml of ethanol (4.1).

Close the tube and shake vigorously. If necessary, place the tube in a water bath heated at 60 °C (5.4) for five minutes in order to completely dissolve the lipid phase. Cool rapidly under running water. Filter part of the solution through a membrane filter (5.5). Chromatograph the filtrate the same day.

6.1.2. Sample preparation with the internal standard

Weigh to three decimal places $1 \pm 0,1$ g (a grams) of the sample into a glass tube (5.3). Add 0,5 ml of phosphoric acid (4.4), 0,50 ml of the internal standard solution (4.6) and 9 ml of ethanol (4.1).

Close the tube and shake vigorously. If necessary, place the tube in a water bath heated to 60 °C (5.4) for five minutes in order to dissolve the lipid phase. Cool rapidly under running water. Filter part of the solution through a membrane filter (5.5). Chromatograph the filtrate the same day.

6.2. Conditions for gas chromatography

The following operation conditions are recommended:

Column

Type	Stainless steel
Length	2 m
Diameter	1/8 "
Packing	10 % SP TM 1000 (or equivalent) + 1 % H ₃ PO ₄ on Chromosorb WAW 100 to 120 mesh

Temperature

Injector	200 °C
Column	120 °C
Detector	200 °C
Carrier gas	nitrogen
Flow rate	25 ml/min

6.3. Chromatography**6.3.1. Calibration**

Into a series of 20-ml volumetric flasks, transfer by pipette 0,25, 0,50, 1,00, 2,00 and 4,00 ml respectively of the propionic acid solution (4.5). To each volumetric flask transfer by pipette 1,00 ml of the internal standard solution (4.6); make up to volume with ethanol (4.1) and mix. The solutions prepared in this way contain e mg/ml of 2-methylpropionic acid as internal standard (that is to say, 1 mg/ml if e = 1 000) and p/4, p/2, p, 2p, 4p mg/ml propionic acid (that is to say, 0,25, 0,50, 1,00, 2,00, 4,00 mg/ml if p = 1 000).

Inject 1 µl of each of these solutions and obtain the calibration curve by plotting the ratio of the propionic acid/2-methylpropionic acid mass on the x-axis and the ratio of the corresponding peak areas on the y-axis.

Make three injections of each solution and calculate the average peak area ratio.

6.3.2. Determination

Inject 1 μ l of the sample filtrate 6.1.1. Compare the chromatogram with that of one of the standard solutions (6.3.1). If a peak has approximately the same retention time as 2-methylpropionic acid, change the internal standard. If no interference is observed, inject 1 μ l of the sample filtrate 6.1.2 and measure the areas of the propionic acid peak and the internal standard peak.

Make three injections of each solution and calculate the average peak area ratio.

7. Calculations

7.1. From the calibration curve obtained in 6.3.1, obtain the ratio of mass (K) corresponding to the peak area ratio calculated in 6.3.2.

7.2. From the ratio of mass thus obtained calculate the propionic acid content of the sample (X) as a percentage by mass using the formula:

$$x \% (m/m) = K \frac{0,5 \cdot 100 \cdot e}{50 \cdot a} = K \frac{e}{a}$$

in which:

K = the ratio calculated in 7.1,

e = mass in grams of the internal standard weighed in 4.6,

a = mass in grams of the sample weighed in 6.1.2.

Round off results to one decimal place.

8. Repeatability ⁽¹⁾

For a propionic acid content of 2% (m/m) the difference between the results of two determinations in parallel carried out on the same sample should not exceed 0,12%.

II. IDENTIFICATION AND DETERMINATION OF HYDROQUINONE, HYDROQUINONE MONOMETHYLETHER, HYDROQUINONE MONOETHYLETHER AND HYDROQUINONE MONOBENZYLETHER IN COSMETIC PRODUCTS

A. IDENTIFICATION

1. Scope and field of application

The method describes the detection and identification of hydroquinone, hydroquinone monomethylether, hydroquinone monoethylether and hydroquinone monobenzylether (monobenzone) in cosmetic products for lightening the skin.

2. Principle

Hydroquinone and its ethers are identified by thin layer chromatography (TLC).

3. Reagents

All reagents must be of analytical grade.

⁽¹⁾ ISO 5725.

- 3.1. Ethanol, 96 % (v/v)
- 3.2. Chloroform
- 3.3. Diethyl ether
- 3.4. Developing solvent:
Chloroform/Diethyl ether, 66:33 (v/v)
- 3.5. Ammonia, 25 % (m/m) ($d_4^{20} = 0,91$ g/ml)
- 3.6. Ascorbic acid
- 3.7. Hydroquinone
- 3.8. Hydroquinone monomethylether
- 3.9. Hydroquinone monoethylether
- 3.10. Hydroquinone monobenzylether (monobenzene)
- 3.11. Reference solutions
The following reference solutions should be freshly prepared, and are stable for one day.
 - 3.11.1. Weigh 0,05 g hydroquinone (3.7) into a 10-ml graduated test tube. Add 0,250 g of ascorbic acid (3.6) and 5 ml of ethanol (3.1). Add ammonia (3.5) until the pH is 10 and make up to a volume of 10 ml with ethanol (3.1).
 - 3.11.2. Weigh 0,05 g hydroquinone monomethylether (3.8) into a 10-ml graduated test tube. Add 0,250 g of ascorbic acid (3.6) and 5 ml of ethanol (3.1). Add ammonia (3.5) until the pH is 10 and make up to a volume of 10ml with ethanol (3.1).
 - 3.11.3. Weigh 0,05 g hydroquinone monoethylether (3.9) into a 10-ml graduated test tube. Add 0,250 g of ascorbic acid (3.6) and 5 ml of ethanol (3.1). Add ammonia (3.5) until the pH is 10 and make up to a volume of 10ml with ethanol (3.1).
 - 3.11.4. Weigh 0,05 g hydroquinone monobenzylether (3.10) into a 10-ml graduated test tube. Add 0,250 g of ascorbic acid (3.6) and 5 ml of ethanol (3.1). Add ammonia (3.5) until the pH is 10 and make up to a volume of 10ml with ethanol (3.1).
- 3.12. Silver nitrate
- 3.13. 12-Molybdophosphoric acid
- 3.14. Potassium ferricyanide hexahydrate
- 3.15. Ferric chloride, hexahydrate
- 3.16. Spray reagents
 - 3.16.1. To a 5 % (m/v) aqueous solution of silver nitrate (3.12), add ammonia (3.5) until the precipitate that forms dissolves.
Warning:
the solution becomes explosively unstable on standing and should be discarded after use.
 - 3.16.2. 10 % (m/v) solution of 12-molybdophosphoric acid (3.13) in ethanol (3.1).

- 3.16.3. Prepare a 1 % (m/v) aqueous solution of potassium ferricyanide (3.14) and a 2 % (m/v) aqueous solution of ferric chloride (3.15). Mix equal parts of both solutions just before use.

4. Apparatus

Normal laboratory equipment, and:

- 4.1. Usual TLC equipment
- 4.2. TLC plates, ready for use: silica gel GHR/UV₂₅₄; 20 × 20 cm (Machery, Nagel, or equivalent). Layer thickness 0,25 mm
- 4.3. Ultrasonic bath
- 4.4. Centrifuge
- 4.5. UV lamp, 254 nm

5. Procedure

5.1. Preparation of the sample

Weigh 3,0 g of sample into a 10-ml graduated tube. Add 0,250 g of ascorbic acid (3.6) and 5 ml of ethanol (3.1). Adjust the pH of the solution to 10, using ammonia (3.5). Make up to a volume of 10 ml with ethanol (3.1). Close the tube with a stopper and homogenize in an ultrasonic bath for 10 minutes. Filter through a filter paper or centrifuge at 3 000 rpm.

5.2. TLC

- 5.2.1. Saturate a chromatographic tank with developing solvent (3.4).
- 5.2.2. Deposit on a plate 2 μ l of the reference solutions (3.11) and 2 μ l of the sample solution (5.1). Develop in the dark at ambient temperature until the solvent front has migrated 15 cm from the start.
- 5.2.3. Remove the plate and allow to dry at room temperature.

5.3. Detection

- 5.3.1. Observe the plate under UV light at 254 nm, and mark the position of the spots.
- 5.3.2. Spray the plate with:
- silver nitrate reagent (3.16.1), or
 - 12-molybdophosphoric acid reagent (3.16.2); heat to approximately 120 °C, or
 - potassium ferricyanide solution and ferric chloride solution (3.16.3).

6. Identification

Calculate the R_f value for each spot.

Compare the spots obtained for the sample solution with those for the reference solutions with respect to: their R_f values; the colour of the spots under UV radiation; and the colours of the spots after visualization with the spray reagent.

Perform the HPLC described in the following section (B), and compare the retention times obtained for the sample peak(s) with those for the reference solutions.

Combine the results from TLC and HPLC to identify the presence of hydroquinone and/or its ethers.

7. Remarks

Under the conditions described, the following R_f values were observed:

hydroquinone:	0,32
hydroquinone monomethylether:	0,53
hydroquinone monoethylether:	0,55
hydroquinone monobenzylether:	0,58

B. DETERMINATION

1. Scope and field of application

This method specifies a procedure for the determination of hydroquinone, hydroquinone monomethylether, hydroquinone monoethylether and hydroquinone monobenzylether in cosmetic products for lightening the skin.

2. Principle

The sample is extracted with a water/methanol mixture under gentle heating to melt any lipid material. Determination of the analytes in the resulting solution is performed by reversed phase liquid chromatography with UV detection.

3. Reagents

- 3.1. All reagents must be of analytical quality. Water used must be distilled water, or water of at least equivalent purity
- 3.2. Methanol
- 3.3. Hydroquinone
- 3.4. Hydroquinone monomethylether
- 3.5. Hydroquinone monoethylether
- 3.6. Hydroquinone monobenzylether (monobenzone)
- 3.7. Tetrahydrofuran, HPLC grade
- 3.8. Water/methanol mixture 1:1 (v/v). Mix one volume of water and one volume of methanol (3.2)
- 3.9. Mobile phase: tetrahydrofuran/water mixture 45:55 (v/v). Mix 45 volumes of tetrahydrofuran (3.7) and 55 volumes of water
- 3.10. Reference solution

Weigh 0,06 g of hydroquinone (3.3), 0,08 g hydroquinone monomethylether (3.4), 0,10 g hydroquinone monoethylether (3.5) and 0,12 g hydroquinone monobenzylether (3.6) into a 50-ml volumetric flask. Dissolve and make up to volume with methanol (3.2). Prepare the reference solution by diluting 10,00 ml of this solution to 50,00 ml with water/methanol mixture (3.8). These solutions must be freshly prepared.

4. Apparatus

Normal laboratory equipment and:

- 4.1. Water bath, capable of maintaining a temperature of 60 °C
- 4.2. High-performance liquid chromatograph with a variable-wavelength UV detector and 10- μ l injection loop
- 4.3. Analytical column:

Stainless steel chromatographic column, length 250 mm, internal diameter 4,6 mm, packed with Zorbax phenyl (chemically bonded phenethylsilane on Zorbax SIL, end-capped with

trimethylchlorosilane), particle size 6 μm , or equivalent. Do not use a guard column, except phenyl guard, or equivalent

- 4.4. Filter paper, diameter 90 mm, Schleicher and Schull, Weissband No 5892, or equivalent

5. Procedure

5.1. Sample preparation

Weigh to three decimal places $1 \pm 0,1$ g (a gram) of sample into a 50-ml volumetric flask. Disperse the sample in 25 ml water/methanol mixture (3.8). Close the flask and shake vigorously until a homogeneous suspension is obtained. Shake for at least one minute. Place the flask in a water bath (4.1) kept at 60 °C to enhance the extraction. Cool the flask, and make up to volume with water/methanol (3.8). Filter the extract using a filter paper (4.4). Perform the HPLC determination within 24 hours of preparing the extract.

5.2. High performance liquid chromatography.

- 5.2.1. Adjust the flow rate of the mobile phase (3.9) to 1,0 ml/min and set the detector wavelength to 295 nm.

- 5.2.2. Inject 10 μl of the sample solution obtained as described in section 5.1, and record the chromatogram. Measure the peak areas. Perform a calibration as described under 5.2.3. Compare the chromatograms obtained for sample and standard solutions. Use the peak areas and the response factors (RF) calculated under 5.2.3 to calculate the concentration of the analytes in the sample solution.

5.2.3. Calibration

Inject 10 μl of the reference solution (3.10) and record the chromatogram. Inject several times until a constant peak area is obtained.

Determine the response factor RF_i :

$$\text{RF}_i = \frac{p_i}{c_i}$$

in which:

p_i = peak area for hydroquinone, hydroquinone monomethylether, hydroquinone monoethylether or hydroquinone monobenzylether, and

c_i = concentration (g/50 ml) in the reference solution (3.10) of hydroquinone, hydroquinone monomethylether, hydroquinone monoethylether or hydroquinone monobenzylether.

Ascertain whether the chromatograms obtained for a standard solution and the sample solution meet the following requirements:

- the peak separation of the worst separated pair shall be at least 0,90. (For definition of peak separation, see Figure 1.)

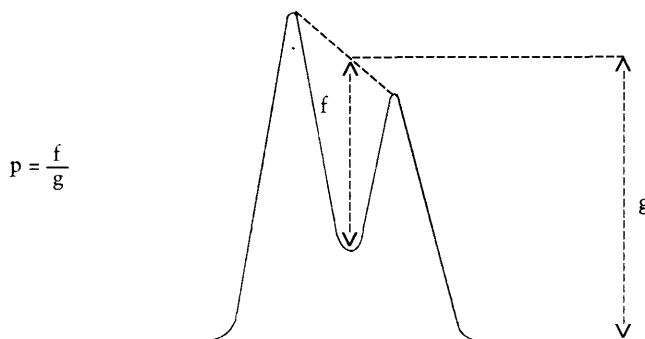


Figure 1: Peak separation

If the required separation is not achieved, either a more efficient column should be used, or the mobile phase composition should be adjusted until the requirement is met.

- The asymmetry factor A_s , of all peaks obtained shall range between 0,9 to 1,5. (For definition of the peak asymmetry factor, see Figure 2.) To record the chromatogram for the determination of the asymmetry factor a chart speed of at least 2 cm/min is recommended.

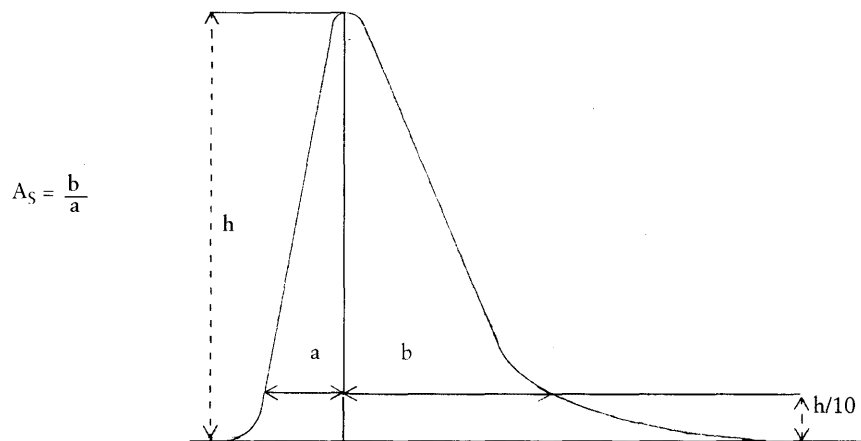


Figure 2: Peak asymmetry factor

- A steady baseline shall be obtained.

6. Calculation

Use the areas of the analyte peaks to calculate the concentration(s) of the analyte(s) in the sample. Calculate the analyte concentration in the sample, as a percentage by mass, (x_i) using the formula:

$$x_i \text{ \% (m/m)} = \frac{b_i \cdot 100}{RF_i \cdot a}$$

in which:

- a = mass of the sample in grams, and
- b_i = peak area of analyte i in the sample.

7. Repeatability ⁽¹⁾

- 7.1. For a hydroquinone content of 2,0% the difference between the results of two determinations carried out in parallel on the same sample should not exceed an absolute value of 0,13%.
- 7.2. For a hydroquinone monomethylether content of 1,0% the difference between the results of two determinations carried out in parallel on the same sample should not exceed an absolute value of 0,1%.
- 7.3. For a hydroquinone monoethylether content of 1,0% the difference between the results of two determinations carried out in parallel on the same sample should not exceed an absolute value of 0,11%.
- 7.4. For a hydroquinone monobenzylether content of 1,0% the difference between the results of two determinations carried out in parallel on the same sample should not exceed an absolute value of 0,11%.

8. Reproducibility ⁽¹⁾

- 8.1. For a hydroquinone content of 2,0% the difference between the results of two determinations carried out on the same sample under different conditions (different laboratories, different operators, different apparatus and/or different time) should not exceed an absolute value of 0,37%.

⁽¹⁾ ISO 5725.

- 8.2. For a hydroquinone monomethylether content of 1,0 % the difference between the results of two determinations carried out on the same sample under different conditions (different laboratories, different operators, different apparatus and/or different time) should not exceed an absolute value of 0,21 %.
- 8.3. For a hydroquinone monoethylether content of 1,0 % the difference between the results of two determinations carried out on the same sample under different conditions (different laboratories, different operators, different apparatus and/or different time) should not exceed an absolute value of 0,19 %.
- 8.4. For a hydroquinone monobenzylether content of 1,0 % the difference between the results of two determinations carried out on the same sample under different conditions (different laboratories, different operators, different apparatus and/or different time) should not exceed an absolute value of 0,11 %.

9. **Remarks**

- 9.1. When a hydroquinone content considerably higher than 2 % is found and an accurate estimate of the content is required, the sample extract (5.1) should be diluted to a similar concentration as would be obtained from a sample containing 2 % hydroquinone, and the determination repeated.

(In some instruments the absorbance may be out of the linear range of the detector for high hydroquinone concentrations.)

9.2. **Interferences**

The method described above allows the determination of hydroquinone and its ethers in a single isocratic run. The use of the phenyl column assures sufficient retention for hydroquinone, which cannot be guaranteed when a C18 column is used with the mobile phase described.

However, this method is prone to interferences by a number of parabens. In such cases the determination should be repeated employing a different mobile phase/stationary phase system. Suitable methods may be found in references 1 and 2, namely:

Column: Zorbax ODS, 4,6 mm × 25 mm, or equivalent:

temperaure: 36 °C

flow: 1,5 ml/min

mobile phase:

for hydroquinone: methanol/water 5/95 (V/V)

for hydroquinone monomethylether: methanol/water 30/70 (V/V)

for hydroquinone monobenzylether: methanol/water 80/20 (V/V) ⁽¹⁾.

Column: Spherisorb S5-ODS, or equivalent:

mobile phase: water/methanol 90/10 (V/V)

flow: 1,5 ml/min.

These conditions are suitable for hydroquinone ⁽²⁾.

⁽¹⁾ M. Herpol-Borremans et M.-O. Masse, Identification et dosage de l'hydroquinone et de ses éthers méthylique et benzylique dans les produits cosmétiques pour blanchir la peau. *Int. j. Cosmet. Sci.* 8-203-214 (1986).

⁽²⁾ J. Firth and I. Rix, Determination of hydroquinone in skin toning creams, *Analyst* (1986), 111, p. 129.